



Cambridge International Examinations

Cambridge International Advanced Subsidiary and Advanced Level

CANDIDATE NAME								
CENTRE NUMBER					CANDIDATE NUMBER			
CHEMISTRY							970	1/33
Paper 3 Advance	ced Prac	tical Skills	s 1		Oc	tober/Nove	mber :	2015
							2 h	ours
Candidates answ	wer on th	e Questi	on Paper.					
Additional Mater	dditional Materials: As listed in the Confidential Instructions							

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Give details of the practical session and laboratory where appropriate, in the boxes provided.

Write in dark blue or black pen.

You may use an HB pencil for any diagrams or graphs.

Do not use staples, paper clips, glue or correction fluid.

DO NOT WRITE IN ANY BARCODES.

Answer all questions.

Electronic calculators may be used.

You may lose marks if you do not show your working or if you do not use appropriate units.

Use of a Data Booklet is unnecessary.

Qualitative Analysis Notes are printed on pages 10 and 11.

A Periodic Table is printed on page 12.

At the end of the examination, fasten all your work securely together. The number of marks is given in brackets [] at the end of each question or part question.

Session
Laboratory

For Examiner's Use	
1	
2	
3	
Total	

This document consists of 12 printed pages.



1 The formula of hydrated copper(II) sulfate is CuSO₄.**x**H₂O where **x** is the number of moles of water of crystallisation in one mole of salt. You will determine the value of **x** by titration.

When aqueous copper(II) ions react with aqueous iodide ions, I^- , iodine is produced.

$$2Cu^{2+}(aq) + 4I^{-}(aq) \rightarrow I_2(aq) + 2CuI(s)$$

The amount of iodine, I_2 , produced can be found by titrating it with aqueous thiosulfate ions, $S_2O_3^{2-}$.

$$2S_2O_3^{2-}(aq) + I_2(aq) \rightarrow S_4O_6^{2-}(aq) + 2I^{-}(aq)$$

FA 1 is aqueous CuSO₄.**x**H₂O containing 26.2 g dm⁻³.

FA 2 is 0.100 mol dm⁻³ sodium thiosulfate, Na₂S₂O₃.

FA 3 is aqueous potassium iodide, KI. starch indicator

(a) Method

- Pipette 25.0 cm³ of **FA 1** into a conical flask.
- Use the measuring cylinder to add 15 cm³ of **FA 3**, an excess of KI, to the conical flask. The solution will turn brown because iodine is formed.
- Fill the burette with FA 2.
- Add **FA 2** from the burette until the colour of the mixture changes to pale brown.
- Add 10 drops of starch indicator. The mixture will turn blue-black.
- Continue adding **FA 2** from the burette until the dark colour suddenly disappears to leave an off-white solid. This is the end point of the titration.
- Carry out a rough titration and record your burette readings in the space below.

The rough	titre is	·	cm ³ .

- Carry out as many accurate titrations as you think necessary to obtain consistent results.
- Make certain any recorded results show the precision of your practical work.

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• Record in a suitable form below, all your burette readings and the volume of **FA 2** added in each accurate titration.

I
II
III
IV
V
VI
VII

(b)		om your accurate titration results, obtain a suitable ow clearly how you have obtained this value.	value to be used in your calculations.
			ired cm ³ of FA 2 . [1]
(c)		Iculations	
	Sho	ow your working and appropriate significant figures	in each step of your calculations.
	(i)	Calculate the number of moles of thiosulfate ions calculated in (b) .	present in the volume of FA 2 you have
		m	noles of S ₂ O ₃ ²⁻ = mo
	(ii)	Use your answer to (i), and the equations for the re of moles of Cu ²⁺ present in 25.0 cm ³ of FA 1 .	eactions involved, to deduce the number
		$2Cu^{2+}(aq) + 4I^{-}(aq) \rightarrow I_2(aq) +$	- 2CuI(s)
		$2S_2O_3^{2-}(aq) + I_2(aq) \rightarrow S_4O_6^{2-}(aq)$) + 2I ⁻ (aq)
			moles of Cu ²⁺ = mo
	(iii)	Use your answer to (ii) and the mass of $CuSO_4$. x H the relative molecular mass, M_r , of $CuSO_4$. x H ₂ O.	H ₂ O present in the solution, to calculate
			$M_{\rm r}$ of CuSO ₄ . \mathbf{x} H ₂ O =
	(iv)	Determine the value of x . (Use data from the Periodic Table on page 12.)	

x =[6]

[Total: 14]

2 FA 4 is an impure sample of hydrated calcium chloride, $CaCl_2.2H_2O$. On heating, hydrated calcium chloride loses its water of crystallisation.

$$CaCl_2.2H_2O(s) \rightarrow CaCl_2(s) + 2H_2O(g)$$

You will determine the purity of **FA 4** by measuring the loss in mass that occurs when it is heated. The impurity present in **FA 4** is not decomposed on heating.

(a) Method

You should read the instructions carefully before starting any practical work and draw a table for your results in the space below.

- Weigh a crucible and record its mass.
- Add between 1.80 g and 2.00 g of **FA 4** into the crucible.
- Reweigh the crucible and its contents and record the mass.
- Place the crucible on the pipe-clay triangle and heat gently for 1 minute and then strongly for a further 2 minutes.
- Allow the crucible and its contents to cool. Reweigh the crucible and contents and record the mass.
- Heat the crucible strongly for a further 2 minutes. Allow it to cool. Reweigh the crucible and contents and record the mass.
- Repeat the heating, cooling and weighing until you are satisfied that all the water of crystallisation has been removed.
- Calculate and record the mass of **FA 4** used and the total mass of water lost.

While you are waiting for the crucible to cool, you may wish to start work on Question 3.

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VI	

[6]

((b)	Ca	lcul	lati	ons
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Show your working and appropriate significant figures in the final answer to **each** stage of your calculations.

(i)	The percentage loss in mass on heating is defined as
	$\frac{\text{the loss in mass on heating}}{\text{the original mass}} \times 100.$
	Calculate the percentage loss in mass of FA 4 .

	percentage loss in mass = %
(ii)	Calculate the percentage loss in mass when ${\bf pure}$ hydrated calcium chloride, ${\rm CaC} l_2.{\rm 2H_2O},$ is heated.
	percentage loss in mass = %
(iii)	Use your results to (i) and (ii) to calculate the percentage purity of FA 4, impure $CaCl_2.2H_2O$.
	percentage purity = %

(c) A student carried out this experiment using 2.60 g of FA 4.

of **FA 4**. Explain your answer.

Suggest whether this experiment would give a more accurate result for the percentage purity

[3]

(d)	In your calculations you assumed that the impurity in FA 4 does not decompose on heating.
	State how the percentage purity that you calculated in (b)(iii) would change if the impurity were to decompose on heating. Explain your answer.
	[1]
	[Total: 11]

3 Qualitative Analysis

At each stage of any test you are to record details of the following.

- colour changes seen
- the formation of any precipitate
- the solubility of such precipitates in an excess of the reagent added

Where gases are released they should be identified by a test, **described in the appropriate place in your observations**.

You should indicate clearly at what stage in a test a change occurs. **No additional tests for ions present should be attempted.**

If any solution is warmed, a boiling tube MUST be used.

Rinse and reuse test-tubes and boiling tubes where possible.

Where reagents are selected for use in a test, the name or correct formula of the element or compound must be given.

(a) (i) FA 5, FA 6 and FA 7 are aqueous solutions each containing one anion and one cation.

Carry out the experiments described below and record your observations for each solution in the table.

	FA 5	FA 6	FA 7
To a 1 cm depth in a test-tube, add a 1 cm depth of aqueous sodium carbonate.			
To a 1 cm depth in a test-tube, add a 1 cm depth of aqueous copper(II) sulfate.			
To a 1 cm depth in a test-tube, add a 1 cm depth of aqueous barium chloride or aqueous barium nitrate.			

(ii)	What ion is present in both FA 6 and FA 7 ?
(iii)	The anion in FA 5 is one of carbonate, chloride, nitrate or sulfate. Which anion is present in FA 5 ?
(iv)	Write the ionic equation, including state symbols, for the reaction between ${\bf FA~5}$ and aqueous copper(II) sulfate.
	[7]

- (b) FA 8 contains two anions and two cations from the lists on pages 10 and 11.
 - To a 5 cm depth of distilled water in a boiling tube, add all the FA 8.
 - Shake the boiling tube thoroughly for one minute to make sure that no more of the solid will dissolve.
 - Filter the mixture into a clean boiling tube.
 - Place the filter funnel in a conical flask and wash the residue with a little distilled water.
 - Keep both filtrate and residue for tests (i) and (ii) below.

(i) Tests on the filtrate (the solution in the boiling tube)

Carry out the following tests and record your observations in the table below.

test	observations
To a 1 cm depth of the filtrate in a test-tube, add aqueous sodium hydroxide, then	
add aqueous hydrogen peroxide.	

(ii) Tests on the residue

Carry out the following tests and record your observations in the table below.

test	observations
Place the funnel containing the residue into a clean boiling tube. Pour approximately 5 cm³ of dilute nitric acid onto the residue. Collect a 1 cm depth of solution in the boiling tube. Remove the funnel and return it to the conical flask.	
To this solution in the boiling tube, add aqueous sodium hydroxide.	

(iii)	Identify two cations present in FA 8 .
	cations present and
(iv)	Identify one anion present in FA 8 .
	anion present
(v)	Suggest what type of reaction is happening when hydrogen peroxide is added in test (b)(i).
	[8]

[Total: 15]

Qualitative Analysis Notes

Key: [ppt. = precipitate]

1 Reactions of aqueous cations

ion	reac	tion with
ion	NaOH(aq)	NH ₃ (aq)
aluminium, A <i>l</i> ³+(aq)	white ppt. soluble in excess	white ppt. insoluble in excess
ammonium, NH₄⁺(aq)	no ppt. ammonia produced on heating	_
barium, Ba ²⁺ (aq)	no ppt. (if reagents are pure)	no ppt.
calcium, Ca²+(aq)	white ppt. with high [Ca ²⁺ (aq)]	no ppt.
chromium(III), Cr³+(aq)	grey-green ppt. soluble in excess giving dark green solution	grey-green ppt. insoluble in excess
copper(II), Cu ²⁺ (aq)	pale blue ppt. insoluble in excess	blue ppt. soluble in excess giving dark blue solution
iron(II), Fe²+(aq)	green ppt. turning brown on contact with air insoluble in excess	green ppt. turning brown on contact with air insoluble in excess
iron(III), Fe³+(aq)	red-brown ppt. insoluble in excess	red-brown ppt. insoluble in excess
magnesium, Mg²+(aq)	white ppt. insoluble in excess	white ppt. insoluble in excess
manganese(II), Mn²+(aq)	off-white ppt. rapidly turning brown on contact with air insoluble in excess	off-white ppt. rapidly turning brown on contact with air insoluble in excess
zinc, Zn²+(aq)	white ppt. soluble in excess	white ppt. soluble in excess

2 Reactions of anions

ion	reaction
carbonate, CO ₃ ²⁻	CO ₂ liberated by dilute acids
chloride, C <i>l</i> ⁻ (aq)	gives white ppt. with Ag ⁺ (aq) (soluble in NH ₃ (aq))
bromide, Br ⁻ (aq)	gives cream ppt. with Ag ⁺ (aq) (partially soluble in NH ₃ (aq))
iodide, I ⁻ (aq)	gives yellow ppt. with Ag ⁺ (aq) (insoluble in NH ₃ (aq))
nitrate, NO ₃ -(aq)	NH ₃ liberated on heating with OH ⁻ (aq) and A <i>l</i> foil
nitrite, NO ₂ ⁻ (aq)	NH_3 liberated on heating with $OH^-(aq)$ and Al foil; NO liberated by dilute acids (colourless $NO \rightarrow$ (pale) brown NO_2 in air)
sulfate, SO ₄ ²⁻ (aq)	gives white ppt. with Ba ²⁺ (aq) (insoluble in excess dilute strong acids)
sulfite, SO ₃ ²⁻ (aq)	SO ₂ liberated with dilute acids; gives white ppt. with Ba ²⁺ (aq) (soluble in excess dilute strong acids)

3 Tests for gases

gas	test and test result
ammonia, NH ₃	turns damp red litmus paper blue
carbon dioxide, CO ₂	gives a white ppt. with limewater (ppt. dissolves with excess CO ₂)
chlorine, Cl ₂	bleaches damp litmus paper
hydrogen, H ₂	"pops" with a lighted splint
oxygen, O ₂	relights a glowing splint
sulfur dioxide, SO ₂	turns acidified aqueous potassium manganate(VII) from purple to colourless

Fm

Es

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Berkeliu

Curium

Pu Plutoniu

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Ра

Thorium

a = relative atomic massX = atomic symbol

a X

Key

The Periodic Table of the Elements

								Gre	Group								
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							1.0 T Hydrogen										4.0 He Helium 2
6.9 Li Lithium 3 23.0 Na Sodium	9.0 Be Beryllium 4 24.3 Mg Magnesium					-						10.8 B Boron 5 27.0 Alt Aluminium 13	12.0 C carbon 6 28.1 Si licon 14	14.0 Nitrogen 7 31.0 Phosphorus 15	16.0 O Oxygen 8 32.1 Sulfur 16	19.0 Fluorine 9 35.5 C.1 Chlorine	20.2 Neon 10 39.9 Argon
39.1 K Potassium	40.1 Ca Calcium	Scandium	47.9 Ti Titanium	50.9 V Vanadium	52.0 Cr Chromium 24	54.9 Mn Manganese 25	55.8 Fe Iron	58.9 Co Cobalt 27	58.7 Ni Nickel 28	63.5 Cu Copper 29	65.4 Zn Zinc 30	69.7 Ga Gallium 31	72.6 Ge Germanium 32	74.9 As Arsenic 33	79.0 Se Selenium 34	79.9 Br Bromine 35	83.8 Kr Krypton 36
Rb Rubidium	87.6 Strontium	88.9 ×	91.2 Zr Zirconium 40	92.9 Nb Niobium 41	95.9 Molybdenum 42	Tc Technetium 43	Ru Ruthenium 44	Rhodium 45	106 Pd Palladium 46	108 Ag Silver 47	Cd Cadmium 48	115 In Indium 49	Sn Tin	122 Sb Antimony 51	128 Te Tellurium 52	127 T lodine	131 Xe Xenon 54
133 Cs Caesium 55	137 Ba Barium 56	139 La Lanthanum 57 *	178 Hf Hafnium 72	181 Ta Tantalum 73	184 W Tungsten 74	186 Re Rhenium	190 Os Osmium 76	192 Ir Iridium	195 Pt Platinum 78	197 Au Gold	Hg Mercury 80	204 T 1 Thallium	207 Pb Lead 82	209 Bi Bismuth 83	Po Polonium 84	At Astatine 85	Radon 86
Francium 87	Radium 88	Actinium 1	Rutherfordium	Db Dubnium 105	Sg Seaborgium 106	Bh Bohrium 107	Hsssium 108	Mt Meitnerium 109	Uun Ununnilium 110	Uuu Unununium 111	Uub Ununbium 112		Uuq Ununquadium 114		Uuh Ununhexium 116		Uuo Ununoctium 118
*58-71 L 190-103	*58-71 Lanthanides 190-103 Actinides	Š	*	140 Cer ium 58	Pr Prassodymium 59	144 Nd Neodymium 60	Pm Promethium 61	150 Sm Samarium 62	152 Eu Europium 63	157 Gd Gadolinium 64	159 Tb Terbium	163 Dy Dysprosium 66	165 Ho Holmium 67	167 Er Erbium	169 Tm Thullum	773 YB Yterbium 70	175 Lu Lutetium 771

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